# **Lesson Plan**

# MCM DAV College for Women, Sector – 36A, Chandigarh Monthly Teaching Plans (Odd Semester) Session: 2023-24

## Name of the Teacher/s: Dr. Shilpa

## **Department:** Chemistry

**Class:** B.Sc-II (3<sup>rd</sup> semester)

#### **Subject:** Physical Chemistry

S.No.	(Mo	Date onthly)	Topics to be Covered	Academic Activity Undertaken*
1	<b>From</b> 21-07-2023	<u>To</u> 04-08-2023	Unit-I: Liquid State Intermolecular forces, structure of liquids (a qualitative description). Structural differences between solids, liquids and gases.	Lecture and group discussion
2	05-08-2023	18-08-2023	Inquites and gases.Unit-I: Liquid StateLiquid Crystals: Differencebetween liquid crystal, solid andliquid. Classification, structureof nematic and cholestric phases.Thermography and sevensegment cell.UNIT-II: Chemical EquilibriumEquilibrium constant and freeenergy.Thermodynamicderivation of law of mass ofmass action. Le - Chatelier'sprinciple.	Lecture and group discussion
3	19-08-2023	27-08-2023	Unit-II Reaction isotherm and Reaction isochore-Clapeyron equation and Clausius –Clapeyron equation, applications. Unit-III: Thermodynamics-II Second Law of Thermodynamics: Need for the law, different statements of the law, Carnot cycle and its efficiency, Carnot theorem. Thermodynamic scale of temperature.	Lecture and group discussion

4	28-08-2	023	15-09-2023	Unit-III	Lecture and group
-	20-00-2	.025	15-07-2025	Concept of Entropy: Entropy as	discussion
				a state function, entropy as a	discussion
				function of V & T, entropy as a	
				function of P & T, entropy	
				change in physical change,	
				Clausius inequality, entropy as a	
				criteria of spontaneity and	
				equilibrium. Entropy change in	
				ideal gases and mixing of gases.	
5	16-09-2	023	24-09-2023	Unit-IV: Thermodynamics-III	Lecture and group
5	10-09-2	.025	24-09-2023	Third Law of Thermodynamics:	discussion
				Nernst heat theorem, statement.	uiscussion
6	25-09-2	023	04-10-2023	Concept of residual entropy,	Lacture and group
U	25-07-2	.023	04-10-2023	evaluation of absolute entropy,	Lecture and group discussion
				from heat capacity data.	u15Cu551011
7	05-10-2	023	20-10-2023	Gibbs and Helmholtz functions;	
/	03-10-2	.025	20-10-2023	Gibbs function (G) and	
				Helmholtz functions (A) as	
				thermodynamic quantities, A &	
				G as criteria for thermodynamic	
				equilibrium and spontaneity,	
				their advantage over entropy	
				0 10	
8	21-10-2	022	Till exams	change. Variation of G and A with P, V	
0	21-10-2	.025	1 III exams		
Doport	and T. and Revision   Departmental Meeting to Coordinate and Review the Monthly completion of Syllabus as per				of Syllabug og nor
Depart		eeting		lesson plans	of Synabus as per
30 <sup>th</sup> Augu	ust. 2023	The t		leted the scheduled chapters and top	pics as shown in the
	.,		n plan		
Depar	rtmental N			nd Review the Monthly completio	on of Syllabus as per
lesson pla			0		· I
29 <sup>th</sup> Sep		The t	eachers have comp	leted the scheduled chapters and to	pics as shown in the
1 /			n plan	1	
Depart	mental M	eeting	to Coordinate an	d Review the Monthly completion	of Syllabus as per
-				lesson plans	• •
31 <sup>st</sup> Oct, 2023		The teachers have completed the scheduled chapters and topics as shown in the			
			n plan	1	
Depart	mental M	eeting		d Review the Monthly completion lesson plans	of Syllabus as per
22 <sup>nd</sup> N	ov, 2023	The t		leted the scheduled chapters and to	pics as shown in the
10	, 2020		n plan	and sensories enuptors and top	
*Any of the	(i)			(iii) Online Sources: (iv) Group Disc	ussion:

\*Any of these – (i) Lecture Method; (ii) PPT; (iii) Online Sources; (iv) Group Discussion; (v) Case Studies etc. Other Methods adopted by the teacher – Please write the specific teaching method

## Lesson Plan

# MCM DAV College for Women, Sector – 36A, Chandigarh Monthly Teaching Plans (Even Semester) Session – (2023-24)

#### Name of the Teacher/s: Dr. Shilpa

#### **Department:** Chemistry

Class: B.Sc-II (4<sup>th</sup> Semester)

## Subject: Physical Chemistry

S.No.	Da (Mon		Topics to be Covered	Academic Activity
	From	То		Undertaken*
1	09 -01- 2024	31-01-2024	Unit-I: Phase equilibrium: Statement and meaning of the terms – phase, component and degree of freedom, derivation of Gibbs phase rule, phase equilibria of one component system—water, CO2 and S systems. Phase equilibria of two component system—solid—liquid equilibria, simple eutectic – Bi-Cd system, desiliverisation of lead. Solid solutions—compound formation with congruent melting point (Mg-Zn) and incongruent melting point, (NaCl- H2O) system. Freezing mixtures, acetone-dry ice.	Lecture and Group Discussion
2	01.02.2024	29.02.2024	Unit-I Partially Miscible Liquids –Phenol- water, trimethylamine – water, nicotine –water systems. Nernst distribution law-thermodynamic derivation, applications. Unit-II: Electrochemistry –I Electrical transport –Conduction in metals and in electrolyte solutions, specific conductance and equivalent conductance, measurement of equivalent conductance, variation of equivalent and specific conductance with dilution. Migration of ions and Kohlrausch Law, Arrhennius theory of electrolyte dissociation and its	Lecture and Group Discussion

			limitations, weak and strong electrolytes, Ostwald's dilution law, its uses and limitations. Debye-Huckel- Onsager's equation for strong electrolytes (elementary treatment only). Transport number, definition and determination by Hittorf method and moving boundary method.	
3	01.03.2024	20.03.2024	Unit-III: Electrochemistry-II Types of reversible electrodes – gas metal – ion, metal –insoluble salt – anion and redox electrodes. Electrode reactions, Nernst equation, derivation of cell E.M.F. and single electrode potential, standard hydrogen electrode – reference electrodes – standard electrode potential, sign conventions, electrochemical series and its significance.	Lecture and Group Discussion
4	21-03-2024	15-04-2024	Unit-IV: Electrolytic and Galvanic cells – reversible and irreversible cells, conventional representation of electrochemical cells. E.M.F. of a cell and its measurements. Computation of cell E.M.F. Calculation of thermodynamic quantities of cell reactions ( $\Delta G$ , $\Delta H$ and K), Polarization, over potential and hydrogen overvoltage. Concentration cell with and without transport, liquid junction potential, application of concentration cells, valency of ions, solubility product and activity coefficient, potentiometric titrations.	Lecture and Group Discussion
5	16.04.2024	Till exam	Previous question papers and doubts sessions	Lecture and Group Discussion

Departmen	Departmental Meeting to Coordinate and Review the Monthly completion of Syllabus as per		
	lesson plans		
30-01-2024	The teachers have completed the scheduled chapters and topics as shown in the lesson		
	plan		
Departmen	Departmental Meeting to Coordinate and Review the Monthly completion of Syllabus as per		
	lesson plans		
24-02-2024	The teachers have completed the scheduled chapters and topics as shown in the lesson		
	plan		
Departmen	Departmental Meeting to Coordinate and Review the Monthly completion of Syllabus as per		
lesson plans			

28-03-2024	The teachers have completed the scheduled chapters and topics as shown in the lesson			
	plan			
Departme	Departmental Meeting to Coordinate and Review the Monthly completion of Syllabus as per			
	lesson plans			
19-04-2024	The teachers have completed the scheduled chapters and topics as shown in the lesson			
	plan			

\*Any of these – (i) Lecture Method; (ii) PPT; (iii) Online Sources; (iv) Group Discussion; (v) Case Studies etc.Other Methods adopted by the teacher – Please write the specific teaching method